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M. Marks: 246 Note: (i)  
The question paper  
consists of 3 parts (Part

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I : Mathematics, Part II  
: Physics, Part III :  
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Question paper format  
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(i) The question paper  
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I : Mathematics, Part II  
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Chemistry). Each part  
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Section I contains 6 multiple choice questions. Each question has 4 choices (A), (B), (C) and (D), out of which only one is correct.

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Question paper format:

1. The question paper consists of 3 parts (Part I: Mathematics, Part II:

Physics, Part III:

Chemistry. Each part has 4 sections. 2.

Section I contains 6

multiple choice

questions. Each

question has 4 choices

(A), (B), (C) and

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lengthier. So although cutoffs will dip this year don't expect them to dip too much. You can check the solutions to IIT JEE 2008.

## **IITJEE 2008 Cutoff, Analysis & Expected Rank ...**

IIT - JEE 2008  
(SOLUTION PAPER -1)  
MATHEMATICS Part I  
Section - I Straight  
Objective Type This  
section contains 6  
multiply choice

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questions. Each question has 4 choices (A), (B), (C) and (D), out of which Only one is correct. 1. Consider the two curves 2  
 $C_1: y = 4x^2$   
 $C_2: x + y - 6x + 1 = 0$  Then,  
(A)  $C_1$  and  $C_2$  touch each other only at one point

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The freezing point of  
the solution M is (A)  
268.7 K (B) 268.5 K  
Sol. (D) Let total moles  
be 1, then nethanol =  
0.9 mass of ethanol =  
41.4 g =  $41.4 \times 10^{-3}$   
kg nwater = 0.1 IIT-JEE  
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(SOLUTION PAPER -1)

MATHEMATICS Part I

Section - I Straight

Objective Type This

section contains 6

multiply choice

questions. Each

question has 4 choices

(A), (B), (C) and (D),

out of which Only one

is correct. 1. Consider

the two curves  $2C_1y$

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$x_1: 4 = 2 \cdot 2$  C  $x \cdot y \cdot x^2: 6 \cdot 1$   
 $0 + - + =$  Then, (A) C1  
and C2 touch each  
other only at one point

## **SOLUTION OF IIT-JEE 2008 (PAPER - I) - Momentum Academy**

JEE Main Question  
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## Solutions

### PAPER 1 CHEMISTRY

Q.1 2.5 mL of 5.2 M weak monoacidic base ( $K_b = 1 \times 10^{-12}$  at  $25^\circ\text{C}$ ) is titrated with 15.2 M HCl in water at  $25^\circ\text{C}$ . The

concentration of  $\text{H}^+$  at equivalence point is ...

A solution M is prepared by mixing ethanol and water. The mole fraction of ethanol in the mixture is 0.9.

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